

TITLE OF THE INVENTION

RECHARGEABLE LITHIUM BATTERY

BACKGROUND OF THE INVENTION**FIELD OF THE INVENTION**

5 The present invention relates to rechargeable lithium batteries and more particularly to rechargeable lithium batteries which use a lithium alloying metal for both or either of the active anode and cathode materials.

DESCRIPTION OF RELATED ART

10 Rechargeable lithium batteries, because of their high energy densities, have been noted to lead the next generation. The use of metallic lithium for the negative active material, although effective to increase a charge-discharge capacity, is accompanied by lithium dissolution
15 and deposition during charging and recharging, dendrite formation at a negative electrode and electrode deformation which together cause poor battery cycle performances and thus prevent provision of practicable batteries.

20 As solutions to such problems, an Li alloy negative electrode using an Li alloying metal and a carbon negative electrode using a carbon material have been proposed. The carbon negative electrode has been put into practical use for years but suffers from a problem of a larger energy density drop compared to the metallic lithium negative
25 electrode, due to its low theoretical capacity, 372 mAh/g.

On the other hand, the Li alloy negative electrode is subjected to increasing pulverization with charge-discharge cycling and thus suffers from a problem of poor cycle performances.

5 For example, Japanese Patent Laying-Open No. Hei 10-312804 proposes a method for improving cycle performances by using a homogeneous single-phase alloy derived via formation of a metal powder by a roll quenching process and subsequent heat treatment thereof.

10 However, the metal powder obtained in accordance with this method consists of large-size, homogenous single-phase alloy particles which are susceptible to a large stress, during the Li alloy formation, that problematically causes active material to fall off from a current collector. Other 15 problem has been that the metal particles tend to grow into treelike dendrites to render active material inactive. As such, those methods have even failed to obtain sufficient cycle performance characteristics.

SUMMARY OF THE INVENTION

20 It is an object of the present invention to provide a rechargeable lithium battery which uses an Li alloying metal as the active material of at least one of positive and negative electrodes and which exhibit excellent cycle performance characteristics.

25 The rechargeable lithium battery of the present

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invention includes a positive electrode, a negative electrode and a nonaqueous electrolyte. Characteristically, an Li alloying metal is used as the active material of at least one of the positive and negative electrodes and this metal active material is covered with a thin film which is non-reactive with Li ions, permits passage of Li ions but does not have an Li ion conductivity.

In the present invention, the provision of such a thin film over the metal active material regulates dendrite formation or pulverization of the metal active material during charge-discharge cycles and thus prevents the metal active material from falling off from a current collector.

The thin film provided to overlie the metal active material in accordance with the present invention is the one that is nonreactive with Li ions, permits passage of Li ions but does not have an Li ion conductivity, as stated above. Because of its nonreactivity with Li ions, the thin film itself is not alloyed and thus avoids the occurrence of expansion or shrinkage. Also, the passage of Li ions across the thin film permits a cell reaction to occur and proceed in metal active material. Also, the thin film itself does not undergo deformation during charging and discharging because it does not have an Li ion conductivity, as contrary to solid electrolyte thin films.

Also in the present invention, the thin film preferably

has a volume resistivity of not exceeding 10^{10} $\Omega\cdot\text{cm}$. Such a good electronic conductivity enables the thin film to also serve as a current collector.

In the present invention, the thin film can be formed
5 such as by a CVD, sputtering or vacuum deposition process.

The thin film in the present invention is illustrated
10 by a hard carbon thin film such as a diamond-like carbon thin film. Such a thin film is the one that does not react with Li ions, permits passage of Li ions and does not have an Li ion conductivity. The preferred hard carbon thin film shows two peaks I_d and I_g in the Raman scattering spectrum, around 1400 cm^{-1} and 1550 cm^{-1} , with a ratio (I_d/I_g) in intensity of 0.5 to 3.0.

As also described above, the preferred hard carbon thin
15 film has a volume resistivity of not exceeding 10^{10} $\Omega\cdot\text{cm}$ and can be illustrated by a CO_2 -containing hard carbon thin film with a good electrical conductivity. The CO_2 -containing hard carbon thin film can be formed by a CVD process using a gas mixture of CO_2 and hydrocarbon as a source gas. Also, a
20 good electrical conductivity can be imparted to a surface of a hard carbon thin film by treating the surface with a CO_2 -containing gas.

The thickness of the thin film is not particularly specified, but is preferably in the approximate range of 50
25 to 1,000 nm, more preferably in the approximate range of 100

DETAILED DESCRIPTION

to 500 nm. If the thin film is too small in dimension, its presence may in some cases become ineffective to prevent the metal active material from falling off from a current collector and to regulate the treelike growth of the metal active material. On the other hand, if the thin film is too large in dimension, the passage of Li ions during charging and discharging may sometimes be regulated to result in the insufficient reaction between the metal active material and Li ions.

As noted earlier, an Li alloying metal serves as the metal active material in the present invention. The Li alloying metal may be at least one metal selected from Si, Ge, Sn, Al, In and Mg. The metal active material may be rendered into a powder form and combined with a binder to form a mix which is then coated on a current collector to fabricate an electrode, in accordance with a general procedure utilized to fabricate an electrode for use in rechargeable lithium batteries. Preferably, the metal active material is used in the form of a thin film. The metal active material may be deposited in the form of a thin film by various techniques, including CVD, sputtering, vapor deposition and plating techniques. In such a case, the metal active material, if deposited on a substrate comprised of a current collector such as a copper foil, can be made into a ready-to-use electrode. Alternatively, a metal foil

may be used as the film-form metal active material. In this case, the aforementioned thin film may be placed on both sides of the film-form metal active material. In the case where a metal foil is used as the film-form metal active material, the metal foil may be allowed to serve as a current collector.

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In the present invention, an interlayer may be provided between the aforesaid thin film and metal active material for purposes including improvement of adhesion therebetween. Such an interlayer may be formed from at least one selected from Si, Ti, Zr, Ge, Ru, Mo, W and their oxides, nitrides and carbides. Preferably, the interlayer has a thickness in the approximate range of 10 to 500 nm. The interlayer can be deposited by various techniques, including CVD, sputtering, vapor deposition and plating techniques.

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The metal active material in the present invention may serve as either a positive or negative active material, but probably in most cases as a negative active material due to its electric potential relative to Li.

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If the latter is the case, any substance conventionally known to serve as positive active material of rechargeable lithium batteries can be used to constitute the positive active material. Examples of useful positive active materials include metal oxides containing at least one metal selected from manganese, cobalt, nickel, vanadium and

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niobium. Examples of specific metal oxides include lithium-containing transition metal oxides such as LiCoO_2 , LiNiO_2 , LiMn_2O_4 , LiMnO_2 , $\text{LiCo}_{0.5}\text{Ni}_{0.5}\text{O}_2$ and $\text{LiNi}_{0.7}\text{Co}_{0.2}\text{Mn}_{0.1}\text{O}_2$; lithium-free metal oxides such as MnO_2 ; and the like. Other substances can also be used, without limitation, if they are capable of electrochemical insertion release of lithium.

The electrolyte solvent for use in the rechargeable battery of the present invention is not particularly limited in type but can be exemplified by a mixed solvent which contains cyclic carbonate such as ethylene carbonate, propylene carbonate or butylene carbonate and also contains chain carbonate such as dimethyl carbonate, methyl ethyl carbonate or diethyl carbonate. Also applicable is a mixed solvent of the above-listed cyclic carbonate and an ether solvent such as 1,2-dimethoxyethane or 1,2-diethoxyethane.

Examples of electrolyte solutes include LiPF_6 , LiBF_4 , LiCF_3SO_3 , $\text{LiN}(\text{CF}_3\text{SO}_2)_2$, $\text{LiN}(\text{C}_2\text{F}_5\text{SO}_2)_2$, $\text{LiN}(\text{CF}_3\text{SO}_2)(\text{C}_4\text{F}_9\text{SO}_2)$, $\text{LiC}(\text{CF}_3\text{SO}_2)_3$, $\text{LiC}(\text{C}_2\text{F}_5\text{SO}_2)_3$ and mixtures thereof. Other applicable electrolytes include, for example, a gelled polymer electrolyte comprised of an electrolyte solution impregnated into a polymer electrolyte such as polyethylene oxide or polyacrylonitrile and inorganic solid electrolytes such as LiI and Li_3N . The electrolyte for the rechargeable lithium battery of the present invention can be used without limitation, so long as an Li compound as its solute that

imparts an ionic conductivity, as well as its solvent that dissolves and retains the Li compound, remain undecomposed at voltages applied during charge, discharge and storage of the battery.

5 **BRIEF DESCRIPTION OF THE DRAWINGS**

Figure 1 is a schematic sectional view which shows a rechargeable lithium battery constructed in Example 1 in accordance with the present invention.

10 **DESCRIPTION OF THE PREFERRED EXAMPLES**

The present invention is now described in more detail with reference to preferred examples. It will be recognized that the following examples merely illustrate the practice of the present invention but are not intended to be limiting thereof. Suitable changes and modifications can be effected without departing from the scope of the present invention.

15 **EXAMPLE 1**

Fabrication of Positive Electrode

20 85 % by weight of LiCoO_2 powder with a mean particle diameter of 10 μm , 10 % by weight of carbon powder as an electrically conductive agent, and 5 % by weight of polyvinylidene fluoride powder as a binder were mixed. N-methyl pyrrolidone was added to the mixture which was then kneaded to form a slurry. The slurry was coated by a doctor blade technique onto both sides of a 20 μm thick current collector made of aluminum (Al). The subsequent drying

resulted in a positive electrode.

Fabrication of Negative Electrode

An Si thin film (20 nm thick) as an interlayer was deposited on each side of an 18 μm thick Sn foil as metal active material by an RF sputtering technique using a target RF power of 200 W and a substrate bias voltage of -100 V. This interlayer was provided to improve adhesion between the Sn foil and a hard carbon thin film which, when directly combined, showed poor adhesion to each other.

Next, a hard carbon thin film (diamond-like carbon thin film) was deposited on the Si thin film by a CVD technique to a thickness of 50, 100, 200, 500, 800 or 1,000 nm. Specifically, an ion beam CVD technique was utilized to deposit the hard carbon thin films under the following thin film-forming conditions; CH_4 gas flow rate of 40 sccm, CO_2 gas flow rate of 10 sccm, ECR plasma source microwave power of 330 W and ion beam acceleration voltage of 200 V.

The resulting hard carbon thin films having different film thicknesses showed the same volume resistivity of about $10^9 \Omega\cdot\text{cm}$. Each of them was measured for Raman scattering spectrum. The Raman scattering revealed a ratio I_d/I_g of about 1.1 for each hard carbon thin film, where I_d = peak strength around $1,400 \text{ cm}^{-1}$ and I_g = peak strength around $1,550 \text{ cm}^{-1}$.

Preparation of Electrolyte Solution

An electrolyte solution was prepared by allowing 1 mole/l of LiPF₆ to dissolve in a mixed solvent containing ethylene carbonate and diethyl carbonate in proportions by volume of 1:1.

5 Construction of Rechargeable Lithium Battery
A rechargeable lithium battery was constructed using the above-obtained positive electrode, negative electrode and electrolyte solution. Figure 1 is a schematic sectional view, illustrating a construction of the rechargeable lithium battery. The rechargeable lithium battery includes a positive electrode 1, a negative electrode 2, a separator 3, a positive case 4, a negative case 5 and an insulating gasket 6 made of polypropylene. A microporous membrane of polyethylene is used for the separator 3. The negative electrode 2 is made by depositing a hard carbon thin film 2b on both sides of an Sn foil 2a through an Si thin film.

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The positive electrode 1 and negative electrode 2 are disposed on opposite sides of the separator 3. These are accommodated within a casing defined by the positive and negative cases 4 and 5. The positive electrode 1 and negative electrode 2 are electrically connected to the positive case 4 and negative case 5, respectively, so that the battery is placed in conditions for ready charge and discharge as a rechargeable battery.

25 The batteries A-1 to A-6 in accordance with the present

invention were constructed having respective thicknesses specified in Table 1. For comparison, a comparative battery B-1 was constructed using, for the negative electrode, an Sn foil on which neither a hard carbon thin film nor an Si thin film as an interlayer was deposited.

EVALUATION OF CHARGE-DISCHARGE CYCLE CHARACTERISTICS

Each of the above-constructed batteries was charged at 20 mA at 25 °C to 4.1 V and discharged at 20 mA to 2.75 V. This unit cycle was repeated 10 times in total. For evaluation of cycle characteristics, a cycle efficiency (%) was calculated from the equation as defined below. The cycle efficiency of each battery is shown in Table 1. Each battery showed a mean discharge voltage of about 3.4 V.

$$\text{Cycle efficiency (\%)} = C_1/C_2 \times 100$$

where,

C_1 = 10th cycle discharge capacity; and

C_2 = 1st cycle discharge capacity.

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[TABLE 1]

| Designation of Battery | Film Thickness (nm) | Cycle Efficiency (%) |
|------------------------|---------------------|----------------------|
| B1 | 0 | 22 |
| A1 | 50 | 60 |
| A2 | 100 | 75 |
| A3 | 200 | 78 |
| A4 | 500 | 77 |
| A5 | 800 | 69 |
| A6 | 1000 | 60 |

As apparent from Table 1, the batteries A-1 to A-6 of the present invention each including a negative electrode fabricated by depositing a hard carbon thin film on an Sn foil (filmed metal active material) through an Si thin film (interlayer) exhibit the improved cycle characteristics compared to the comparative battery B-1 using a negative electrode devoid of a hard carbon thin film.

Each battery after 10 cycles was disassembled to observe a condition of the Sn foil incorporated as negative active material. Observations of the batteries A-1 to A-6 of the present invention revealed that their Sn foils were slightly made into powder but retained their original shapes before charge-discharge cycling. On the other hand, an

observation of the comparative battery B-1 revealed that its Sn foil was completely made into powder.

These observation results demonstrate that the practice of the present invention limits the negative active material 5 from being made into powder and thereby improves cycle performance characteristics.

In accordance with the present invention, improved cycle characteristics can be imparted to rechargeable lithium batteries by the use of an Li alloying metal as active material for at least one of their positive and 10 negative electrodes.